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APPLICATION THAT MET THE REQUIREMENTS TO BE GRANTED A  
FILING DATE.

APPLICATION NUMBER: 60/510,983

FILING DATE: *October 14, 2003*

RELATED PCT APPLICATION NUMBER: PCT/US04/33915

Certified by



Jon W Dudas

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**PROVISIONAL APPLICATION FOR PATENT COVER SHEET**

This is a request for filing a PROVISIONAL APPLICATION FOR PATENT under 37 CFR 1.53(c).

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2258  
60

INVENTOR(S)		
Given Name (first and middle [if any])  J. Donald Jose I.	Family Name or Surname  Carruthers Arno	Residence (City and either State or Foreign Country)  Fairfield, CT Brookfield, CT
<input type="checkbox"/> Additional inventors are being named on the _____ separately numbered sheets attached hereto		
TITLE OF THE INVENTION (500 characters max)  H2 GENERATION		
<p>Direct all correspondence to:</p> <p><input checked="" type="checkbox"/> Customer Number <b>25559</b> → <input type="checkbox"/> Place Customer Number Bar Code Label here</p> <p>OR <input type="checkbox"/> Type Customer Number here</p>		
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ENCLOSED APPLICATION PARTS (check all that apply)		
<input checked="" type="checkbox"/> Specification Number of Pages  <b>11</b>	<input type="checkbox"/> CD(s), Number  <b> </b>	
<input checked="" type="checkbox"/> Drawing(s) Number of Sheets  <b>2</b>	<input type="checkbox"/> Other (specify)  <b> </b>	
<input type="checkbox"/> Application Data Sheet. See 37 CFR 1.76		
METHOD OF PAYMENT OF FILING FEES FOR THIS PROVISIONAL APPLICATION FOR PATENT		
<input type="checkbox"/> Applicant claims small entity status. See 37 CFR 1.27.	FILING FEE	
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Respectfully submitted,

SIGNATURE Date **10/14/2003**TYPED or PRINTED NAME **Margaret Chappuis**REGISTRATION NO.  
(if appropriate)  
Docket Number:**45,735**TELEPHONE **203-794-1100**

ATMI-682-PRV

**USE ONLY FOR FILING A PROVISIONAL APPLICATION FOR PATENT**

This collection of information is required by 37 CFR 1.51. The information is used by the public to file (and by the PTO to process) a provisional application. Confidentiality is governed by 35 U.S.C. 122 and 37 CFR 1.14. This collection is estimated to take 8 hours to complete, including gathering, preparing, and submitting the complete provisional application to the PTO. Time will vary depending upon the individual case. Any comments on the amount of time you require to complete this form and/or suggestions for reducing this burden, should be sent to the Chief Information Officer, U.S. Patent and Trademark Office, U.S. Department of Commerce, Washington, D.C. 20231. DO NOT SEND FEES OR COMPLETED FORMS TO THIS ADDRESS. SEND TO: Box Provisional Application, Assistant Commissioner for Patents, Washington, D.C. 20231.

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# FEE TRANSMITTAL

## for FY 2004

Effective 10/01/2003. Patent fees are subject to annual revision.

 Applicant claims small entity status. See 37 CFR 1.27

TOTAL AMOUNT OF PAYMENT (\$ 160)

## Complete if Known

Application Number	
Filing Date	
First Named Inventor	J. Donald Carruthers
Examiner Name	
Art Unit	
Attorney Docket No.	ATMI-682-PRV

## METHOD OF PAYMENT (check all that apply)

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 Deposit Account:Deposit Account Number  
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## FEE CALCULATION

## 1. BASIC FILING FEE

Large Entity	Small Entity	Fee Code (\$)	Fee Code (\$)	Fee Description	Fee Paid
1001 770	2001 385	Utility filing fee			
1002 340	2002 170	Design filing fee			
1003 530	2003 265	Plant filing fee			
1004 770	2004 385	Reissue filing fee			
1005 160	2005 80	Provisional filing fee	160		
SUBTOTAL (1)		(\$ 160)			

## 2. EXTRA CLAIM FEES FOR UTILITY AND REISSUE

Total Claims	Independent Claims	Extra Claims	Fee from below	Fee Paid
		-20** =	X	=
		-3** =	X	=

Large Entity	Small Entity	Fee Description
1202 18	2202 9	Claims in excess of 20
1201 86	2201 43	Independent claims in excess of 3
1203 290	2203 145	Multiple dependent claim, if not paid
1204 86	2204 43	** Reissue independent claims over original patent
1205 18	2205 9	** Reissue claims in excess of 20 and over original patent
SUBTOTAL (2)		(\$)

\*\*or number previously paid, if greater. For Reissues, see above

## FEE CALCULATION (continued)

## 3. ADDITIONAL FEES

Large Entity Small Entity

Fee Code (\$)	Fee Code (\$)	Fee Description	Fee Paid
1051 130	2051 65	Surcharge - late filing fee or oath	
1052 50	2052 25	Surcharge - late provisional filing fee or cover sheet	
1053 130	1053 130	Non-English specification	
1812 2,520	1812 2,520	For filing a request for ex parte reexamination	
1804 920*	1804 920*	Requesting publication of SIR prior to Examiner action	
1805 1,840*	1805 1,840*	Requesting publication of SIR after Examiner action	
1251 110	2251 55	Extension for reply within first month	
1252 420	2252 210	Extension for reply within second month	
1253 950	2253 475	Extension for reply within third month	
1254 1,480	2254 740	Extension for reply within fourth month	
1255 2,110	2255 1,005	Extension for reply within fifth month	
1401 330	2401 165	Notice of Appeal	
1402 330	2402 165	Filing a brief in support of an appeal	
1403 290	2403 145	Request for oral hearing	
1451 1,510	1451 1,510	Petition to institute a public use proceeding	
1452 110	2452 55	Petition to revive - unavoidable	
1453 1,330	2453 665	Petition to revive - unintentional	
1501 1,330	2501 665	Utility issue fee (or reissue)	
1502 480	2502 240	Design issue fee	
1503 640	2503 320	Plant issue fee	
1460 130	1460 130	Petitions to the Commissioner	
1807 50	1807 50	Processing fee under 37 CFR 1.17(q)	
1806 180	1806 180	Submission of Information Disclosure Stmt	
8021 40	8021 40	Recording each patent assignment per property (times number of properties)	
1809 770	2809 385	Filing a submission after final rejection (37 CFR 1.129(a))	
1810 770	2810 385	For each additional invention to be examined (37 CFR 1.129(b))	
1801 770	2801 385	Request for Continued Examination (RCE)	
1802 900	1802 900	Request for expedited examination of a design application	
Other fee (specify)			
Reduced by Basic Filing Fee Paid			
SUBTOTAL (3)		(\$)	

(Complete if applicable)

Name (Print/Type)	Margaret Chappuis	Registration No. (Attorney/Agent)	45,735	Telephone	203-794-1100
Signature	Margaret Chappuis				
Date	10/14/2003				

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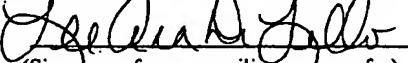
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Docket No.: ATMI-682-PRV  
Customer ID No.: 25559

In re application of: J. Donald Carruthers, et al.

TITLE : H2 GENERATION

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The following papers are enclosed:

Provisional Application for Patent Cover Sheet (1 pg)  
Fee Transmittal (1 pg) (In Duplicate)  
Specification (11)  
2 Sheets of Drawings (Figs. 1, 2 and 6)

**UNITED STATES PROVISIONAL PATENT APPLICATION**

**OF**

**J. DONALD CARRUTHERS**

**AND**

**JOSE I. ARNO**

**FOR**

**H<sub>2</sub> GENERATION**

**EXPRESS MAIL CERTIFICATE OF MAILING**

Express Mail Label Number: EV178871405US  
Date of Deposit: October 14, 2003

## H<sub>2</sub> GENERATION

### Background

In the January 2003 State of the Union Address, President Bush reaffirmed the Nation's commitment to energy independence and an improved environment by proposing \$1.2 billion in research funding for hydrogen and fuel cell technologies over the next five years. While hydrogen production and fuel cell technologies to support this goal are well advanced, the greatest challenge ('Grand Challenge') is undoubtedly the need for a cost-effective, on-board hydrogen storage system that will meet the DOE minimum vehicle range of 300 miles within the weight and volume constraints of the vehicle. DOE emphasizes that this is the greatest challenge, since no hydrogen storage technology available today can meet the DOE cost and performance targets.(1)

Of the technical challenges facing this goal, on-board storage of hydrogen is perhaps the most formidable. Although hydrogen can be stored in several ways, e.g on a solid adsorbent, as a cryogenic liquid, as a compressed gas, or even as a solid chemical hydride, significant barriers must be overcome with each of these methods before the targeted goals can be achieved.

The problem of hydrogen storage is multifaceted. Various approaches have been tried: adsorption of H<sub>2</sub> on inert solids, storing liquid petroleum or methanol followed by reforming to H<sub>2</sub>, and decomposition of solid hydrides. Conventional adsorption methods and materials have been shown to be completely inadequate. Reports in 1997 of high hydrogen adsorption levels on carbon nanotube adsorbents appeared to supply an answer(2, 4), but attempts to repeat the work were disappointing(3, 5). McEnaney reviewed the state of the art in a review paper in 2003(6) and concluded that numerous claims had been made, but there was little convincing evidence that hydrogen can be adsorbed at the levels required. Indeed, should hydrogen adsorb at the levels originally claimed(2), the specific gravity of hydrogen at room temperature within the carbon nanotube would have to exceed that of solid hydrogen by a factor of 2.88(7); quoting the author of that work: "...such a supercompacted form of hydrogen is strange to contemporary physics...". Other theoretical studies arrived at the same conclusion(8).

Finally, at a recent international meeting of carbon scientists (Carbon 2003, Oviedo, Spain), there was an open panel discussion of hydrogen adsorption on carbon nanotubes, in which the question was posed to the audience of 150 people whether anyone had convincing evidence that the levels of hydrogen adsorption necessary to meet DOE targets could be reached. It is critical that alternative approaches be considered as well.

Hydrogen can be stored as a chemically-bonded hydride, and much work is underway at National Laboratories(9) and industrial concerns(10) to demonstrate such technology. This work centers on the use of solid hydrides of the type NaAlH<sub>4</sub> doped with certain activating elements such as Ti. Complex hydrides, of the form Mg(AlH<sub>4</sub>)<sub>2</sub>•AlH<sub>6</sub>(11) are also under investigation. These technologies, while promising, introduce other challenges, such as the necessary change to solids handling and away from fluids handling in the energy distribution system, the poor gravimetric energy density of the fuel,(6) and the fact that the solid hydrides must be heated to relatively high temperatures in order to release hydrogen.

Hydrogen production *via* reforming of methanol and/or methane has received some attention.(12) Adsorbed Natural Gas (ANG) powered vehicles were evaluated extensively in the 1990's,(13) but the use of ANG as a source of hydrogen would require a similar level of technology as proposed for petroleum fuel processing for fuel cell powered vehicles, and would still generate large amounts of CO<sub>2</sub>.

The military has identified the need for small portable electric power supplies. The U.S. infantryman, for example, has become extremely efficient through the use of high tech. devices; e.g. devices which provide him with night vision capability and now I.R. 'vision'. These devices however require increasing amounts of portable electric power. Currently-available battery packs are both heavy and unwieldy and function for only a few hours at a time before requiring recharge. Re-charging devices using fuel cells are under development but these require hydrogen as fuel, supplied either from a compressed gas cylinder or more usually, by catalytic treatment of a liquid fuel such as methanol. One downside to the use of methanol is the fact that the catalytic treatment process uses some portion of the fuel in its conversion to hydrogen thereby reducing its efficiency. This approach is still quite attractive with the foot-soldier carrying refill

cartridges of methanol. The other option of carrying a high pressure, heavy compressed hydrogen gas cylinders is not acceptable.

### Brief Description of the Drawings

Figure 1: Conceptual schematic of hydrogen storage via silane adsorption and regeneration.

Figure 2: Silane decomposition to H<sub>2</sub> & Si on untreated carbon (various tests) and H<sub>3</sub>PO<sub>4</sub>-treated carbon.

Figure 3: Silane adsorption/desorption isotherms at 294.2°K

### Detailed Description

The following United States Patents are incorporated herein in their entirety, by reference in their entirety. United States Patents 5518528 and 6089027.

The invention described here relates to an alternative to conventional hydrogen storage methods. In essence, the invention we propose has three separate stages:

- A.- Safe storage of a hydride material in a reduced pressure package
- B.- Method to decompose the hydride in order to generate H<sub>2</sub>
- C.- Ability to regenerate (*in* or *ex-situ*) the left-over metal back to the hydride

Note that the essential invention is a combination of steps (A) and (B); The last regeneration step (C) is not necessary but closes the materials lifecycle loop. The three stages are described next in addition to systems integration and examples of materials suitable for H<sub>2</sub> generation.

**Safe Hydride Gas Storage:** Many hydride gases are toxic posing environmental or human safety hazards. Therefore, a major barrier for accepting the concept of *in-situ* generation of hydrogen from a hydride precursor are concerns over accidental high pressure gas release. Extensive investigations have been made by ATMI(14) to demonstrate the value of the Safe Delivery System. Tests were conducted by remotely

opening a conventional phosphine cylinder valve located in a secure cabinet and measuring the maximum gas evolution rate during the release period. The release rate of phosphine from the conventional high pressure cylinder attained levels as high as 29,000 ppm/minute. The maximum release rate reached by opening the valve of a SDS™ cylinder was only 480 ppb/minute. These tests and nearly 10 years of perfectly safe performance have convinced the semiconductor industry that these gases can be handled safely in a routine manner. As many as 40,000 SDS™ cylinders are currently in use, worldwide. In addition to the aforementioned safety advantages of storing gases in the adsorbed phase, it has been recently discovered that they can reduce the risk of deflagration when storing deflagration-prone gases (such as germane). Because of the "unstable" nature of some of the hydrides, adsorbents may need to be passivated to prevent premature or uncontrolled decomposition. For example, boric acid treatment is a well-known method for oxidation suppression in carbon materials, (15) where it is believed to involve bonding of  $(BO_3)_n$  to the edge of the {101} face of graphite. We thus believe that the passivating effect observed in our own data is good evidence for the importance of *carbon edge-site chemistry* in hydride decomposition. The following are potential features of the hydride storage system:

- The characteristics of the adsorbent are the same of the ones described in the SDS or SAGE patents. Pore size distribution will be optimized to yield optimum storage and desorption rates. The can be carbon or zeolite based. In addition, the media can be pelletized, beaded or a solid monolith block.
- Another gas storage method that could be suitable for risk reduction in this invention, is the use of a mechanical regulator (such as ATMI's VAC' package) inside the vessel to prevent high-pressure gas release.
- Moreover, additional safety advantages can be achieved using a combination of adsorbent and mechanical regulator technologies.
- The pressure vessel storing the hydride can be sub-atmospheric or at reduced pressures compared to conventional storage but above atmospheric pressures.
- In yet another version of this invention, gas is stored in an adsorbent and a check valve prevents the potentially dangerous situation where air back diffuses to the container.

Hydride Decomposition:

We propose a novel approach to hydrogen storage. A recent discovery at ATMI has demonstrated that silane gas ( $\text{SiH}_4$ ), stored as an adsorbed-phase in a nano-composite medium at sub-atmospheric pressures, can self-decompose at a near constant rate, producing  $\text{H}_2$  and solid silicon. This decomposition occurs at room temperature without the need for external heat or elevated pressure. This exciting, new hydrogen storage concept involves, therefore, the release of hydrogen via the decomposition of a hydride gas stored in a nano-composite, adsorbed-phase vessel according to the reaction:



where  $A$  denotes an element capable of forming a gaseous hydride.

The thermodynamic tendency for this reaction to occur at room temperature is indicated by the Gibbs free energy of formation of the hydride at 298K,  $\Delta G_f^\circ$ . The decomposition of certain hydrides are more thermodynamically favored than others. Stibine ( $\text{SbH}_3$ ,  $\Delta G_f^\circ = +147.8 \text{ kJ/mol}$ ), germane ( $\text{GeH}_4$ ,  $\Delta G_f^\circ = +113.4 \text{ kJ/mol}$ ), diborane ( $\text{B}_2\text{H}_6$ ,  $\Delta G_f^\circ = +86.7 \text{ kJ/mol}$ ), and silane ( $\text{SiH}_4$ ,  $\Delta G_f^\circ = +56.9 \text{ kJ/mol}$ ) are gaseous hydrides which, subject to kinetic limitations, can spontaneously decompose at room temperature. On the other hand, the decomposition of methane ( $\text{CH}_4$ ,  $\Delta G_f^\circ = -50.84 \text{ kJ/mol}$ ) is very unfavorable under these conditions. Of these gases, silane stands out as the leading candidate based on safety and abundance considerations.

Decomposition rates can be modulated using the following methods:

- Heating the vessel to increase kinetic decomposition rates.
- Pressurizing the vessel to increase kinetic rates
- Modify the surface sites of the carbon to catalyze decomposition: There is the possibility of developing higher-rate decomposition catalysts by increasing the number of active edge sites on carbon. Data developed at ATMI showed enhanced decomposition rates after doping with phosphoric acid, implying a synergist mechanism involving carbon edge sites and phosphorous. Overall, we believe that our

ability to systematically manipulate both edge-site concentration and the type, amount, and dispersion of inorganic dopants, there is great potential to find an effective catalyst formulation for rapid, on-demand hydrogen production from the silane source. Figure 1 indicates the decomposition of silane as a function of carbon site functionalization.

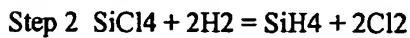
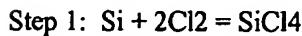
- Use some of the hydrogen being generated to cool or heat up the vessel to slow down or accelerate H<sub>2</sub> generation.

Material Regeneration:

Once the hydride has been fully utilized to generate hydrogen, the non-gaseous metal is left in the vessel. That metal can be recycled back to the hydride form in order to complete the material lifecycle. This can be more effectively accomplished in or ex-situ. Unlike current reforming technologies used to generate hydrogen which liberate the second element into the environment, (usually carbon as carbon oxides), this technology aims to confine the other element (Si) for conversion back to a hydrogen-containing species. From the thermodynamic point of view, the enthalpy required for formation of the silane is +34.3 kJ/mole while the heat of combustion of the generated hydrogen reaches -483.6 kJ, a net overall benefit of -449 kJ. Methods for hydride regeneration include:

- Use hydrogen and heat/and or pressure to convert the metal back to the hydride
- Use another hydride with the ability to donate some of its hydrogen atoms to regenerate the leftover metal. For example, a higher order borane (such as decaborane) could donate some of its hydrogens to react with boron and form diborane. Energy balances will dictate best alternatives.
- Use a halogen gas to extract the metal to generate metal halides; as a second stage, displace the halogen atoms via reaction with H<sub>2</sub> (with heat, catalysts, or pressure) to form the hydride.

For example using silicon,



Integrated System:

In one embodiment of this invention, the hydride would be stored in one vessel and the hydrogen generation modulated by adding or removing energy to the system (heat, pressure, friction, etc). Because the H<sub>2</sub>, by a large extent, would not be preferentially bonded onto the adsorbent media, it will remain at the headspace of the container. Whenever the gas is vented out, most of the headspace gas will be extracted out (namely H<sub>2</sub>). Any hydride accidentally removed together with the gaseous H<sub>2</sub> can be filtered out of the stream using a membrane or chemical filter.

In another embodiment, hydride storage and decomposition occur in two separate medias. The first material would consist of an adsorbent as described in section (A). The second stage, stored in a separate vessel or in a separate partition, will contain a catalyst that decomposes the hydride into a solid and hydrogen gas (as described in section B). As with the previous embodiment, hydrogen gas will be separated from the remaining hydride gas using membranes or filters. The preceding considerations suggest the following conceptual configuration (Figure 2). Hydride gas transfer from storage to decomposition vessels will be controlled using a venturi or mechanical pump.

Membranes will operate based on molecular size or polarity. For instance, Nafion membranes have been known to separate gases based on these characteristics. Check valves pressure switches and/or mechanical regulators can be used throughout the system to prevent backflow and to control generation of hydrogen and to ensure safe operation. In addition, hydrogen sensors can be integrated inside and outside the system to monitor generator performance and for safety purposes.

Materials Proposed:

The hydride materials proposed for this invention must generate hydrogen upon decomposition. Preferably, the by-products should be in the solid phase and regenerable back to their hydride form. Optimum materials should form multiple H<sub>2</sub> molecules for every parent hydride; at the same time, toxicity and reactivity of the hydride and its by-products should be minimized. A final consideration for extensive world-wide use is availability. The following are a few candidates:

For example, silane can be adsorbed at ambient pressure on ATMI, high capacity carbon monoliths at a level of 140g/L. Assuming stoichiometric decomposition, the amount of silane necessary for 4.5% hydrogen storage is 286 g/L. Hence current materials can already achieve approximately 50% of the 2005 DOE hydrogen storage capacity at ambient pressure. The advantages offered by silane (e.g., it is a commodity chemical, less toxic than the other candidate compounds, and capable of safe delivery), far outweigh its disadvantages. Disilane or higher-order silanes can also be used.

Ammonia is readily used in industrial processes. Compared to silane, hydrogen yield is lower but it is less reactive. The by-product is safe N<sub>2</sub> that can be easily separated from hydrogen. Decomposition of ammonia into N<sub>2</sub> and H<sub>2</sub> is not trivial (from the energy point of view) and catalysts may be required.

Boranes (diborane, penta-borane, deca-borane) are great potential sources of H<sub>2</sub>. Upon decomposition, they form boron metal that can be oxidized or re-hydrided.

Hydrocarbons: Readily available and may contain plenty on H atoms. Decomposition to carbon is likely to require a catalyst and energy (similar to ammonia). Regeneration may be difficult if stored in a carbon matrix as the by-product is the same as the adsorbent.. Using a different adsorbent (such as a zeolite) may prove beneficial.

In gas storage operations, one downside is the continual decomposition of the silane or germane gas to hydrogen and silicon/germanium during storage even AT ROOM TEMPERATURE. It now appears that this downside could be turned into upside potential by offering this system as a hydrogen source. The prime novelty of this discovery being its ability to supply hydrogen gas continually at room temperature until the silane is fully decomposed.

Experiments conducted recently reveal that silane decomposition with formation of hydrogen gas can be either suppressed or enhanced depending upon which doping agent is applied to the carbon prior to gas storage (Fig. 1). Even more important is evidence that the decomposition continues, even years after initial storage (Fig. 2).

A 100 gram sample of carbon (perhaps only 100 cc volume) carrying an adsorbed phase of 22 grams of silane could generate 30.8 liters of hydrogen gas when the silane is fully decomposed. Decomposition rates can be adjusted by impregnating the carbon with various dopants. Boric acid dopant, for example, reduces the rates of decomposition dramatically while phosphoric acid doping of the carbon prior to adsorption of the silane gas has been shown to double the rate of hydrogen production over that of the un-treated carbon, all at room temperature.

What is proposed then is a light-weight, small volume, gas storage container rated to ~50 psig, comprising silane gas, SiH<sub>4</sub>, at sub-ambient pressure, adsorbed on an activated carbon which has been doped with a doping agent such as phosphoric acid to accelerate silane decomposition. The gas dispensing mechanism would incorporate a selective hydrogen separation membrane to prevent discharge of silane gas but allow hydrogen flow to a small, portable fuel cell. An infantryman carrying the hydrogen-supply device incorporated into a small fuel cell system would have electric power to continually charge a small battery system. By careful selection of carbon selection, and carbon treatment, hydrogen rates could be designed to match the hydrogen demand of the fuel cell thereby maximizing the life of the hydrogen supply.

When the gas source has been fully discharged, the cylinders could be opened to the atmosphere allowing silicon oxidation to environmentally benign silica before disposal.

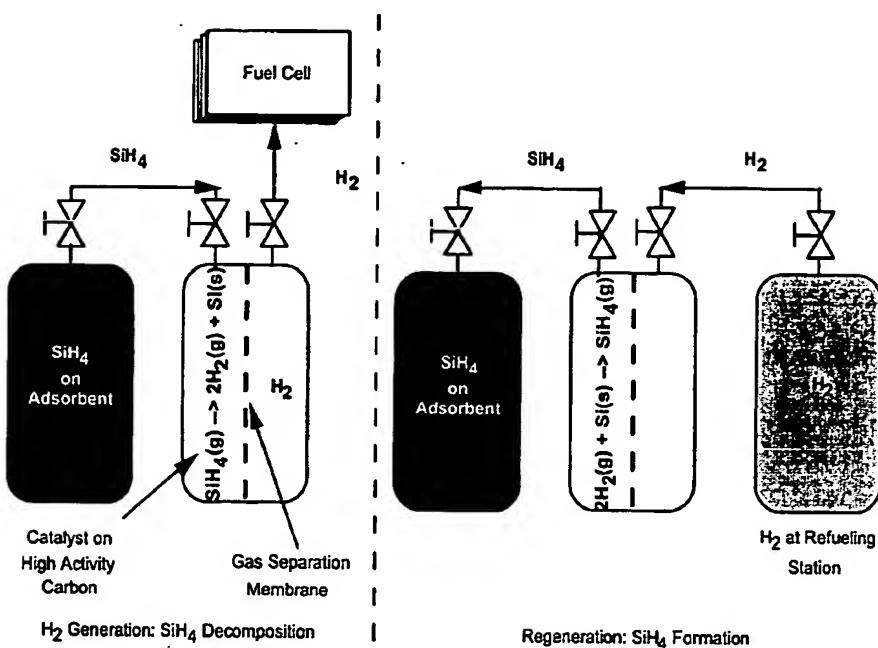
Silane gas is not only toxic but also pyrophoric and its accidental release could be problematic. By supplying the gas at sub-atmospheric pressure these concerns are partially alleviated. Even so, an infantryman would be expected to carry high explosive devices so the additional risk associated with carrying a silane source is probably small.

It is important to add that other metallic hydrides and organometallic hydrides are known to generate hydrogen on storage on a carbon adsorbent so the invention should not be limited to the use of silane and germane only.

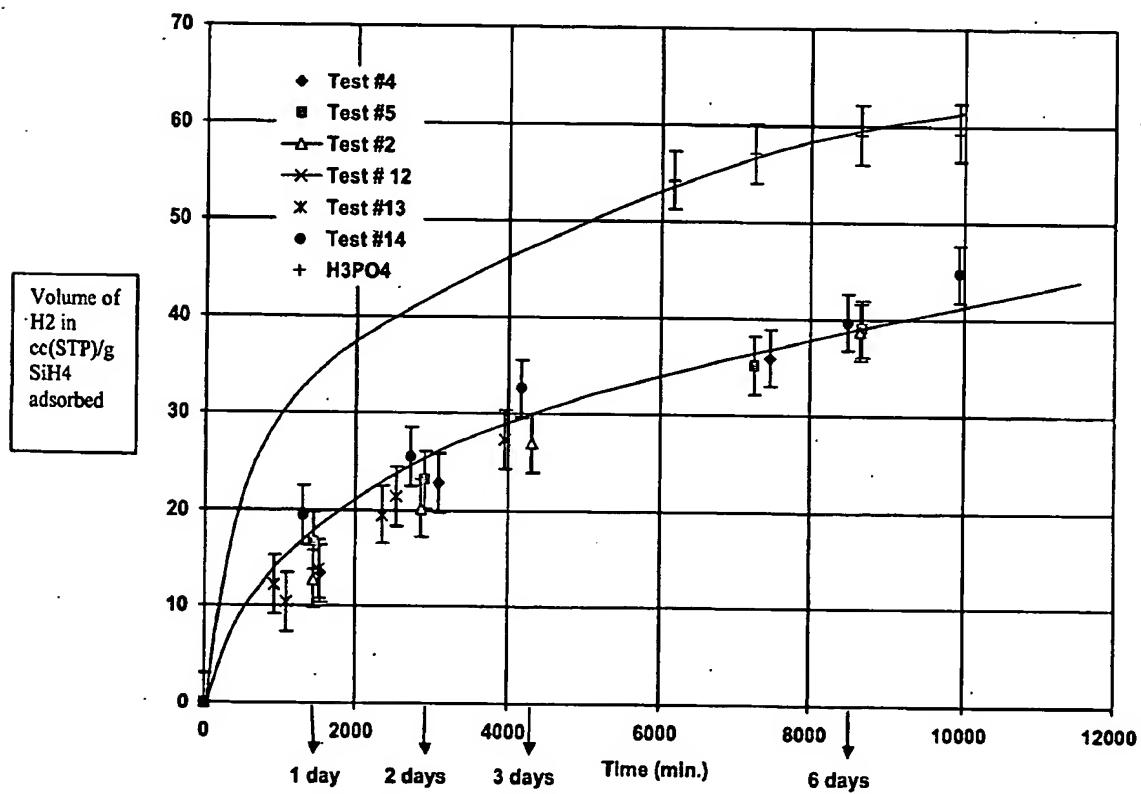
### References Cited

- <sup>1</sup> Farrauto, R., Preprints, ACS Division of Fuel Chemistry, 226<sup>th</sup> ACS National Meeting New York September 2003, paper No. 87 "Catalysts for the Hydrogen Economy".
2. Dillon, A.C., Jones, K.M., Bekkedahl, T.A., Kiang, C.H., Bethune, D.S., Heben, M.J., Nature (London) (1997), 386(6623), 377-379, "Storage of Hydrogen in single-walled carbon nanotubes".
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**Attachment 1. Figures.**



**Figure 1.** Conceptual schematic of hydrogen storage *via* silane adsorption, and regeneration.



**Figure 2.** Silane decomposition to hydrogen and silicon on untreated carbon (various tests) and H<sub>3</sub>PO<sub>4</sub>-treated carbon. Note: this was a study of undesirable decomposition during storage. Much faster rates can be expected using optimized decomposition catalysts.

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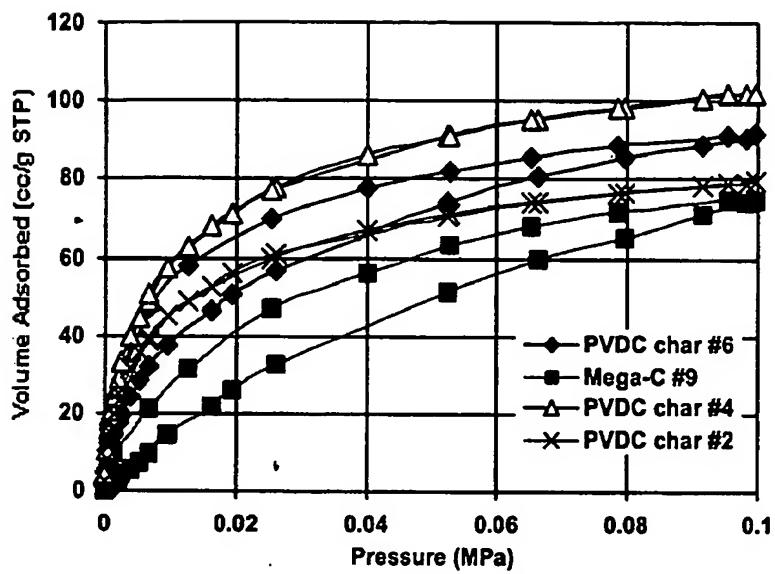


Figure 6. SiH<sub>4</sub> isotherms at 294.2K.

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